Chemistry – pH practice

Calculate the following:

1. What is the **pH** of a solution where:

a) $[H+] = 4.3 \times 10^{-5}$

b) $[OH-] = 8.4 \times 10^{-4}$

c) pOH = 8.3

d) $[H+] = 7.6 \times 10^{-10}$

e) $[OH-] = 2.2 \times 10^{-9}$

f) pOH = 3.7

2. What is the **[H+]** of a solution where:

a) pH = 11.2

b) pOH = 9.4

c) $[OH-] = 6.2 \times 10^{-11}$

d) pH = 3.6

e) pOH = 2.2

f) $[OH-] = 1.8 \times 10^{-5}$

3. What is the pH of a 0.05M HCl solution?

4. What is the pH of a 0.05M NaOH solution?